**☺ C6 6MQ practice ☺**

Answer the questions in 6MQ style. Use the keywords underneath to help your answer.

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| 1. Describe the standard procedure for purifying a salt from solution.   **Evaporation; Filtration; Crystallisation; Dissolving** |
| 1. Explain how to carry out a titration, discussing how to ensure precision of measurements and reliability of results.   **Endpoint; Neutral; Dropwise; Burette; Conical flask; Acid; Alkali; Meniscus** |
| 1. Explain the difference between endothermic and exothermic reactions, and why the chemical industry may wish to manage the energy changes that occur in reactions.   **Energy; Given out; Taken in; Rate of reaction; Cooling** |
| 1. Explain some of the important factors that industrial chemists may alter in order to increase or decrease the rate of an industrial reaction.   **Temperature; Surface area. Size; Concentration; Catalyst; Collision** |
| 1. Explain what neutralisation is, using the example of a reaction between sulphuric acid and calcium hydroxide. Use ionic equations and diagrams to show the reaction and work out the formula of the salt formed.   **Ion; Salt; Water; Acid; Alkali** |